

PRACTICE PROBLEMS - Chemical Formula Writing ws 2

1. Write the formula for the following binary ionic compounds given the name:

- | | |
|----------------------|--------------------------------|
| a. calcium oxide | f. iron (II) sulfide |
| b. potassium bromide | g. silver nitride |
| c. zinc fluoride | h. chromium (III) phosphide |
| d. gold (III) iodide | i. hydrogen sulfide |
| e. aluminum chloride | j. tin (IV) oxide |

2. Write the formula for the following compounds with polyatomic ions given the name:

- | | |
|----------------------------|--------------------------------|
| a. copper (II) nitrate | f. mercury (II) chromate |
| b. barium carbonate | g. ammonium phosphate |
| c. zinc hydrogen phosphate | h. lead (IV) sulfate |
| d. gold (III) hydroxide | i. chromium (II) acetate |
| e. aluminum permanganate | j. phosphorous acid |

3. Write the formula for the following binary molecular compounds:

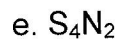
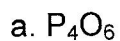
- | | |
|--------------------------------|-----------------------------|
| a. disilicon hexafluoride | f. iodine monochloride |
| b. sulfur hexabromide | g. oxygen difluoride |
| c. sulfur trioxide | h. nitrogen tetrabromide |
| d. tetraphosphorus decasulfide | i. disulfur heptoxide |
| e. diarsenic trioxide | j. phosphorus pentachloride |

4. Write the formula for the compounds given using Parts 1, 2, and 3 above:

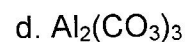
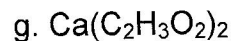
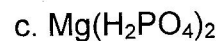
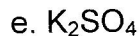
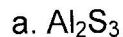
- | | |
|------------------------------------|----------------------------|
| a. lead (II) nitride | h. tin (IV) phosphate |
| b. barium (III) nitrate | i. beryllium nitrite |
| c. triboron octachloride | j. oxygen dichloride |
| d. diphosphorus tetraiodide | k. lead (IV) carbonate |
| e. iodine | l. tetrasulfur hexahydride |
| f. copper (II) hydroxide | m. iron (III) acetate |
| g. triarsenic dinitride | n. strontium phosphate |

PRACTICE PROBLEMS – Nomenclature for Ionic and Molecular Compounds

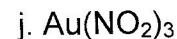
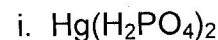
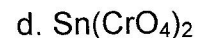
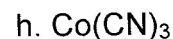
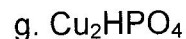
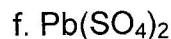
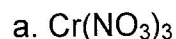
5. Name the following binary molecular compounds:



6. Name the following ionic (fixed charge) compounds:



7. Name the following ionic (variable charge) compounds:



8. Name the following ionic or molecular compounds. Use the Nomenclature Rules

